

1. The 1st to 8th successive ionisation energies, in kJ mol^{-1} , of an element in period 3 are:

1012 1903 2912 4957 6274 21,269 25,398 29,855

What is the element?

- A Al
- B Si
- C P
- D S

Your answer

[1]

2. Which statement best explains why nitrogen has a larger first ionisation energy than oxygen?

- A N atoms have less repulsion between p-orbital electrons than O atoms.
- B N atoms have a smaller nuclear charge than O atoms.
- C N atoms lose an electron from the 2s subshell, while O atoms lose an electron from the 2p subshell.
- D N atoms have an odd number of electrons, while O atoms have an even number.

Your answer

[1]